



METHODS OF WHITENING TEETH

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Related U.S. Application(s)

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The present application is a continuation of United States Application Serial No. 89/192,609, filed on November 16, 1998, itself a divisional of United States Application Serial No. 08/719,569, filed on September 25, 1996, which issued on July 13, 1999 as U.S. Patent No. 5,922,307, and which claims priority from Provisional Application Serial No. 60/004,258, filed September 25, 1995, all of which are hereby incorporated herein by reference.

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Technical Field

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The present invention relates to improved dental bleaching compositions and methods for bleaching teeth.

Background Art

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White teeth have long been considered cosmetically desirable. Unfortunately, teeth become almost invariably discolored in the absence of intervention. The tooth structures which are generally responsible for presenting a stained appearance are enamel, dentin, and the acquired pellicle. Tooth enamel is predominantly formed from inorganic material, mostly in the form of hydroxyapatite crystals and further contains approximately 5% organic material primarily in the form of collagen. In contrast, dentin is composed of about 20% protein including collagen, the balance consisting of inorganic material, predominantly hydroxyapatite crystals, similar to that found in enamel. The acquired pellicle is a proteinaceous layer on the surface of tooth enamel which reforms rapidly after an intensive tooth cleaning.

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Staining of teeth results from extrinsic and/or intrinsic staining. Extrinsic staining of the acquired pellicle arises as a result of compounds such as tannins and polyphenolic compounds which become trapped in and tightly bound to the proteinaceous layer on the surface of the teeth. This type of staining can usually be removed by mechanical methods of tooth cleaning. In contrast, intrinsic staining occurs when staining compounds penetrate the enamel and even the dentin or arise from sources within the tooth. This type of staining is not amenable to mechanical methods of tooth cleaning and chemical methods are required.

Consequently, tooth-bleaching compositions generally fall into two categories: (1) gels, pastes, or liquids, including toothpastes that are mechanically agitated at the stained tooth surface in order to affect tooth stain removal through abrasive erosion of stained acquired pellicle; and (2) gels, pastes, or liquids that accomplish the tooth-bleaching effect by a chemical process while in contact with the stained tooth surface for a specified period, after which the formulation is removed. In some cases, the mechanical process is supplemented by an auxiliary chemical process which may be oxidative or enzymatic.

The majority of professionally-monitored at-home tooth-bleaching compositions act by oxidation. These compositions are dispensed directly to a patient for use in a custom-made tooth-bleaching tray, held in place in the mouth for contact times of greater than about 60 minutes, and sometimes as long as 8 to 12 hours. The slow rate of bleaching is in large part, the consequence of formulations that are developed to maintain stability of the oxidizing composition. The most commonly used oxidative compositions contain the hydrogen peroxide precursor carbamide peroxide which is mixed with an anhydrous or low-water content, hygroscopic viscous carrier containing glycerine and/or propylene glycol and/or polyethylene glycol. When contacted by water, carbamide peroxide dissociates into urea and hydrogen peroxide. Associated with the slow rate of bleaching in the hygroscopic carrier, the currently available tooth-bleaching compositions cause

tooth sensitization in over 50 % of patients. Tooth sensitivity is believed to result from the movement of fluid through the dentinal tubes toward nerve endings in the tooth. This movement is enhanced by the carriers for the carbamide peroxide. In fact, it has been determined that glycerine, propylene glycol and polyethylene glycol
5 can each give rise to varying amounts of tooth sensitivity following exposure of the teeth to heat, cold, overly sweet substances, and other causative agents.

Prolonged exposure of teeth to bleaching compositions, as practiced at present, has a number of adverse effects in addition to that of tooth sensitivity. These include: solubilization of calcium from the enamel layer at a pH less than 5.5
10 with associated demineralization; penetration of the intact enamel and dentin by the bleaching agents, so as to reach the pulp chamber of a vital tooth thereby risking damage to pulpal tissue; and dilution of the bleaching compositions with saliva with resulting leaching from the dental tray and subsequent digestion.

The stability of existing formulations of hydrogen peroxide-containing tooth-
15 bleaching compositions in terms of shelf-life as well as over the period of use in the mouth, depends, in general, on an acidic pH. The hydrogen peroxide becomes markedly less stable as the pH increases. Indeed, Frysh, et al. (Journal of Esthetic Dentistry Vol. 7, No. 3, pp. 130-133, 1995) described the use of high concentration (35%) of hydrogen peroxide solutions at an initial alkaline pH, which was required
20 to be formulated immediately before use and was administered in the form of a liquid to extracted teeth to achieve tooth bleaching. Phillips and Bowles (IADR Abstract J. Dent.res 75, 1996) have demonstrated that hydrogen peroxide penetrates the enamel of extracted teeth less readily over a 15 minute period at pH 9.0 than at pH 4.5. Carbamide peroxide compositions have been formulated at a pH of 5.0-6.5
25 using hygroscopic carriers and maintaining a low water content. This type of formulation is problematic with regard to enhanced tooth sensitivity. On contact with saliva, the water content of the formulation increases, causing the carbamide to disassociate into urea and hydrogen peroxide and the pH to be decreased. In fact,

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Thus, currently available tooth-bleaching compositions that rely on hydrogen peroxide as oxidizing agents, all release hydrogen peroxide from precursors at low pH levels despite the low rates of tooth-bleaching activity.

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baffle in response to the externally applied pressure.

A further embodiment of the invention includes a method for bleaching teeth including preparing a composition as described above and administering the composition to the tooth surface.

Detailed Description of the Invention

The present invention comprises compositions and methods for bleaching tooth enamel *in situ* which allow the use of reduced concentrations of hydrogen peroxide in tooth-bleaching compositions in order to achieve effective tooth bleaching in a contact time of less than one hour.

The tooth surface is defined here and in the claims as a portion of a tooth which is directly responsible for the stained appearance of said tooth. The term tooth surface generally means a tooth's acquired pellicle, plaque, enamel, and combinations thereof.

The matrix is defined here and in the claims as the gel, paste, or liquid in which the hydrogen peroxide containing compound is placed for administration to the subject.

The subject referred to here and in the claims is commonly a human subject but also includes domestic animals.

An important aspect of the present invention is the finding that the efficiency of the bleaching reaction in a tooth using a chemical tooth-bleaching agent such as hydrogen peroxide can be significantly enhanced at a pH greater than 5.5, more particularly a pH in the range of 6-10, for example in a range of pH of 7-10, more particularly between 8.0 and 9.5, providing that the pH is maintained at a substantially constant range throughout the tooth-bleaching process and a calcium chelating agent is included in the composition to prevent precipitation of calcium ions. (Table 1-4)

Suitable pH adjusting agents include, but are not limited to, sodium

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hydroxide, potassium hydroxide, ammonium hydroxide, sodium carbonate, potassium carbonate, TRIS and triethanolamine.

Examples of calcium chelating agents include any of the calcium chelating agents known in the art. Examples include EDTA and its salts, citric acid and its salts, gluconic acid and its salts, alkali metal pyrophosphates and alkali metal polyphosphates. The use of citric acid, sodium acid pyrophosphate and disodium EDTA are shown in Table 1 and Table 2. The biological efficacy of sodium acid pyrophosphate is shown in Table 4. Without being bound to any particular theory, it is proposed that calcium precipitation in the form of calcium phosphates arise in the intercrystalline interstices of the tooth at elevated pHs and this gives rise to a blockage of movement of the peroxide into the tooth with a resulting negative effect on tooth bleaching. Calcium chelating agents may prevent this precipitation of calcium ions with the associated observed improvement of tooth-bleaching effect.

The composition may also contain a stabilizing agent for removing from solution, metal ions that interfere with the action of the hydrogen peroxide. In certain formulations, a single component may act either as a calcium chelating agent or as a stabilizing agent or may serve both functions.

The ability to maintain a constant pH range above 5.5 throughout the period of tooth bleaching coupled with the inhibition of calcium precipitation that normally occurs at pH levels above 5.5, has resulted in a formulation that is suitable for use in the home and is capable of achieving detectable tooth bleaching in less than 30 minutes. Furthermore, multiple treatments (each treatment lasting no longer than 30 minutes) have been demonstrated to substantially whiten teeth (Example 5). In Example 5, 14 treatments using the inventive formulation, were administered for twenty minutes, twice a day over seven days. A mean ΔE of 7.32 was observed denoting a substantial tooth-bleaching effect compared with a ΔE of 4.73 using a prior art formulation for a period of time that was three times longer than that of the novel formulation.

According to the invention, additional agents having tooth-bleaching effect may be used to achieve detectable tooth bleaching in less than 30 minutes. For example, sodium percarbonate has been demonstrated to be very effective at tooth bleaching when maintained at a pH that is greater than 5.5, more preferably in the

5 range of 7-10, more particularly 8-9.5 that includes a calcium chelating agent. This composition differs from carbamide peroxide in that there is no acidification of that solution that results upon its dissociation. Consequently, the reagent may be maintained at a constant elevated pH for an extended period of time without the necessity for adding buffering reagents beyond that naturally supplied in the saliva.

10 Unlike carbamide peroxide, however, the percarbonate is prepared in a formulation that does not include glycerine and is only slowly permeated by water whereupon hydrogen peroxide is released. For this reason, it may be desirable, but is not essential, to prepare the percarbonate in a two-component composition, the two components being mixed before use so as to accelerate the tooth-bleaching effect

15 (Examples 1 and 2). Formulations containing two components may be applied to the dental tray by squeezing a tube in much the same way as a single component. The mixing of the two components can be readily achieved using a multi-component tube containing a baffle, otherwise known in the art as a static mixer such that on squeezing the tube, material from each of the compartments is forced through the

20 static mixer and are mixed together before emerging from a single exit in the tube.

The present invention has important health benefits that follow from shorter contact times of the tooth with hydrogen peroxide as well as the need for lower concentrations of peroxide to achieve a desired tooth-bleaching effect.

25 **Example 1: Composition of a stable tooth-bleaching formulation**
suitable for use in a one component system.

The formulations below utilized ultrapure components to avoid

destabilization caused by metal ion contaminants. The chelating agent used here is one of disodium EDTA (C), citric acid (1B), and sodium acid pyrophosphate (1F). The pH is modified using one of sodium hydroxide monohydrate (1A, 1B, 1C), ammonium hydroxide (1F, 1G), Tris(hydroxymethyl) aminomethane (1D), and
5 triethanolamine (1E). Carbopol is a high molecular weight crosslinked polyacrylic acid thickening agent. Hydrogen peroxide is used as the oxidizing agent.

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Table 1

Example	A	B	C	D	E	F	G
Ingredient	WEIGHT PERCENT						
Distilled Water	86.41	86.21	86.31	72.80	79.52	86.50	73.81
1-Hydroxyethylidene-1,1-diphosphonic acid	0.02	0.02	0.02	0.03	0.02	0.30	0.40
Sodium stannate trihydrate	0.02	0.02	0.02	0.03	0.02	0.05	0.05
Citric acid	—	0.20	—	—	—	—	0.10
Calcium disodium EDTA	—	—	0.10	—	—	—	—
Sodium acid pyrophosphate	—	—	—	—	—	0.30	—
Hydrogen Peroxide 35%	10.30	10.30	10.30	17.14	17.14	8.60	17.14
Carbopol 974P (BF Goodrich)	2.50	2.50	2.50	5.00	—	3.00	5.00
Carbopol 934P (BF Goodrich)	—	—	—	—	2.00	—	—
Sodium Hydroxide Monohydrate	to pH 7.0	to pH 7.0	to pH 7.0	—	—	—	—
Ammonium hydroxide 28%	—	—	—	—	—	to pH 6.5	to pH 8.5
Tris(hydroxymethyl) aminomethane	—	—	—	to pH 8.0	—	—	—
Triethanolamine	—	—	—	—	to pH 6.0	—	—
Total	100	100	100	100	100	100	100
pH @ 25 deg. C	7.0	7.0	7.0	8.0	6.0	6.5	6.5

The above formulations were prepared by dissolving the stabilizers 1-

- 5 hydroxyethylidene-1,1-diphosphonic acid and the sodium stannate trihydrate in the distilled water using a Kynar-coated propeller-type agitator (reserving enough water, if necessary, to dissolve the neutralizer in the final step). The hydrogen peroxide was then added slowly to this mixture. The Carbopol 974P was then added to the distilled water/stabilizer/hydrogen peroxide mixture slowly while a
- 10 vortex was formed with the agitator blade. This dispersion was then placed in a

Kynar-coated vacuum double planetary mixer to which the pH adjusting agent was added slowly to affect neutralization of the Carbopol 974P and to adjust the final composition pH. The resulting composition was a transparent, viscous gel and was packaged in foil/plastic laminate tubes having a polyethylene product contact liner.

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Example 2: A two component alkaline tooth-bleaching formulation.

A hydrogen peroxide-releasing composition was formulated which utilized sodium percarbonate in an anhydrous gel, and was designed to be combined with a separate aqueous gel prior to use in order to dissolve the sodium percarbonate to form hydrogen peroxide and sodium carbonate. The pH of this composition, shown in Table 2, when combined in a volume ratio of 1 to 1, was 9.0. The chelating agents used here are EDTA and sodium acid pyrophosphate.

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Table 2

EXAMPLE 2	PART A	PART B
INGREDIENT	WEIGHT PERCENT	
Distilled Water	—	95.20
Polyethylene glycol 400	83.00	—
Sodium percarbonate (powder)	12.00	—
EDTA	—	0.20
Sodium acid pyrophosphate	—	0.30
Carbopol 974P (BF Goodrich)	2.50	4.00
Sodium Hydroxide Monohydrate	—	to pH 7.0
Tris(hydroxymethyl)aminomethane	2.50	—
Total	100	100

5 The above composition parts were prepared as follows:

Part A - The Carbopol was dispersed in the polyethylene glycol using a propeller-type agitator with Kynar-coated product contact surfaces. The resulting dispersion was added to a Kynar-coated vacuum double planetary mixer (as in
10 Example 1) and neutralized with the tris(hydroxymethyl)aminomethane under low shear mixing. To the resultant neutralized Carbopol gel, sodium percarbonate was added and dispersed until the composition was a homogenous white paste.

Part B - The Carbopol was added to the water (in which the EDTA and
15 sodium acid pyrophosphate had already been dissolved) and agitated as above. The resulting dispersion was transferred to the Kynar-coated vacuum double planetary mixer and neutralized with the sodium hydroxide monohydrate under low shear mixing. The final gel was clear and viscous, with a pH of around 7.0.

Example 3: Preparation of a tooth-bleaching formulation having acidic pH

In order to demonstrate the superior ability of the inventive hydrogen peroxide gel composition of Examples 1 and 2 to bleach teeth, a composition was prepared which was similar to that of Example 1E, except that the pH was adjusted to 4.5.

Table 3

INGREDIENT	PERCENT
Distilled water	79.82
1-Hydroxyethylidene-1,1-diphosphonic acid	0.02
Sodium stannate trihydrate	0.02
Carbopol 974P	2.00
Hydrogen Peroxide 35%	17.14
Triethanolamine	1.00
TOTAL	100.00

The formulation was prepared as in Example 1, resulting in a transparent, viscous gel with a pH of approximately 4.5. The formulation is similar to 1E, the difference lying in the pH of the composition.

Example 4: Assay to determine tooth bleaching

Bovine incisors, which had been imbedded in an acrylic matrix such that the buccal surfaces presented themselves on the top surface, were stained in a manner to duplicate the tooth staining observed in vivo by humans (alternately exposed to air and a staining broth at 37 degrees C. containing typticase soy broth, tea, coffee, mucin, FeCl₃, and *Sarcina lutea*, for a period of about four days). Each stained bovine incisor was numbered and measured for degree of staining (color by the CIELAB protocol) with a Minolta 5031 Spectrophotometer (3mm aperture, 8 exposure

averaging, outliers discarded). Incisors were covered with different tooth-bleaching compositions in the tables above, in addition to a commercially available carbamide peroxide composition (Opalescence 10% Carbamide Peroxide Gel, Ultradent, South Jordan, UT). All gels were kept in contact with the incisor surface for exactly 15 minutes, whereupon the tooth was rinsed clean of any gel residue with distilled water and swabbed with saliva. The degree of stain removal was thereafter immediately determined by measuring the incisor surface, as above, for color, and the change in tooth color recorded below as ΔE . Absolute color change is defined as the square root of the sum of the squares of all color components (L, a, and b).

$$\sqrt{[(\Delta L)^2 + (\Delta a)^2 + (\Delta b)^2]} = \Delta E$$

Table 4

Tooth #	Product/ Example	pH (neat)	Initial Color			Final Color			ΔE
			L	a	b	L	a	b	
1	Opalescence	6.5	41.79	3.17	11.78	44.29	2.96	11.70	2.51
2	Example 3	4.5	39.84	4.99	12.00	43.96	4.47	10.94	4.29
3	1E	6.0	40.44	4.41	9.53	46.32	3.48	7.54	6.27
4	1A	7.0	36.02	3.84	10.10	42.57	2.59	8.28	6.91
5	1B	7.0	38.81	3.98	11.38	45.92	2.38	8.81	7.73
6	1C	7.0	36.90	4.05	12.61	44.11	2.45	10.53	7.67
7	1D	8.0	41.55	3.67	10.51	49.77	1.26	7.82	8.98
8	1F	6.5	38.55	5.01	10.87	44.78	3.67	9.50	6.52
9	1G	8.5	40.26	4.59	9.93	48.28	3.13	7.97	8.38
10	Example 2	9.0	36.49	4.00	12.64	44.93	2.20	10.63	8.78

This table shows the effect of pH on tooth bleaching. As shown for tooth #2 treated with the formulation of Example 3 and tooth #3 treated with the formulation

of 1E in Example 1, the increase in pH from 4.5 (2) to 6.0 (3) results in an increased AE from 4.29 to 6.27.

The table further shows the positive effect of the calcium chelating agent on tooth bleaching. For example, for 1A, 1B, and 1C (all at pH 7.0), 1A lacked a calcium
5 chelating agent whereas 1B and 1C contained a chelating agent. There was an observed improvement in ΔE in the presence of the chelating agent. The best tooth-bleaching results were obtained at the highest pH, namely, in this experiment, pH 8.0 and pH 9.0.

Opalescence is a commercial product which has been pH adjusted to pH 6.5
10 before use but shows a poor performance with regard to color change over the time of the experiment. It is proposed that the pH of the formulation is lowered as hydrogen peroxide and urea is released following dissociation of carbamide peroxide.

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15 **Example 5: In vivo demonstration of tooth bleaching**

Six volunteers aged 25 to 43 were separated into two groups of two and custom dental trays were fashioned for each participant in the study.

One group was given an unmarked 2 oz. tube containing the composition of Example 1B and instructed to place a small amount of tooth-bleaching material into
20 the tray, position the tray over the teeth, and leave the tray in place for 20 minutes. Patients were instructed to repeated this procedure twice a day for one week, for a total of 14 treatments and 280 minutes total tooth whitener exposure time.

The second group was given an unmarked 2 oz. tube of Opalescence 10% Carbamide Peroxide tooth-bleaching gel and instructed as above, with the exception
25 of the duration of the bleaching procedure to be 60 minutes. Patients were instructed to repeat the procedure twice a day for one week, for a total of 14 treatments and 840 minutes total tooth-bleaching exposure time.

The results of direct tooth surface (upper left central incisor) color measurements, both before and after treatment (as in Example 5 above), are recorded in the Table 5 below.

5 Table 5

Patient #	Product/ Example	Treatment Time (minutes)	Initial Color			Final Color			ΔE
			L	a	b	L	a	b	
1	1B	280	53.76	4.65	11.65	60.34	0.97	8.80	8.06
2	1B	280	49.42	2.97	9.48	56.99	0.46	7.38	8.25
3	1B	280	51.26	2.33	8.25	55.63	0.87	4.99	5.65
4	Opalescence	840	52.78	1.75	6.14	57.26	1.42	2.10	6.04
5	Opalescence	840	56.35	1.79	5.21	59.13	0.65	2.44	4.09
6	Opalescence	840	55.71	2.72	7.10	58.60	1.09	4.75	4.07

The average ΔE for the Example 1B group was 7.32, whereas the average ΔE for the Opalescence group was 4.73. The present inventive compositions are thus shown to offer a substantially improved degree of tooth bleaching in a shorter exposure time than a prior art composition.

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